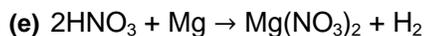
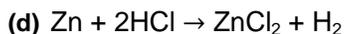
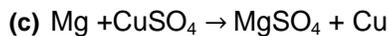
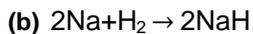
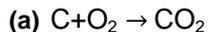


Oxidation and Reduction Questions

1 For each of the following reactions identify the species which is being oxidized and the species which is being reduced.



2 For each of the reactions in question 1, identify the oxidizing agent and the reducing agent.

3 For each of the reactions in question 1, write two half reactions.

4 Write down the oxidation number of the element underlined in the following

a) \underline{H}_2O	b) $\underline{P}Cl_3$
c) $\underline{P}Cl_5$	d) $H_2\underline{S}O_4$
e) $H_2\underline{S}$	f) $\underline{Pb}O_2$
g) $\underline{Pb}O$	h) $\underline{C}Cl_4$
i) $H\underline{N}O_3$	j) \underline{Mg}^{2+}
k) $Mg\underline{S}O_4$	l) $\underline{Mg}CO_3$
m) $\underline{Mn}O_2$	n) $K\underline{N}O_3$
o) $\underline{C}O_2$	p) \underline{O}^{2-}
q) $\underline{Fe}I_3$	r) $\underline{Cu}O$
s) \underline{K}_2O	t) \underline{Br}^-