

C3 – Energetics Quiz

1. How do you know if a reaction is endothermic or exothermic?
2. Is bond breaking exothermic or endothermic?
3. How do you calculate the total energy change from bond energy data?
4. Define activation energy.
5. How does a catalyst increase the rate of reaction?
6. Why does energy need to be supplied at the start of an exothermic reaction but the reaction continues by itself afterwards?
7. How do you prevent energy loss in a colorimetry experiment?
8. How do you measure the energy released when a fuel is burned?
9. A reaction happens rapidly without the help of a catalyst. What does this suggest about the activation energy?
10. Why is hydrogen a good fuel for planes and rockets?
11. Sketch a fully labelled energy profile for an exothermic reaction.
12. Sketch a fully labelled energy profile for an endothermic reaction.
13. How does a hydrogen fuel cell work?