

Please write clearly in block capitals.

Centre number

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Candidate number

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Surname

Model Answers

Forename(s)

Candidate signature

## A-level CHEMISTRY

Paper 2 Organic and Physical Chemistry

Tuesday 12 June 2018

Afternoon

Time allowed: 2 hours

### Materials

For this paper you must have:

- the Periodic Table/Data Booklet, provided as an insert (enclosed)
- a ruler with millimetre measurements
- a scientific calculator, which you are expected to use where appropriate.

### Instructions

- Use black ink or black ball-point pen.
- Fill in the boxes at the top of this page.
- Answer **all** questions.
- You must answer the questions in the spaces provided. Do **not** write outside the box around each page or on the blank pages.
- Do all rough work in this book. Cross through any work you do not want to be marked.

### Information

- The marks for questions are shown in brackets.
- The maximum mark for this paper is 105.

| For Examiner's Use |      |
|--------------------|------|
| Question           | Mark |
| 1                  |      |
| 2                  |      |
| 3                  |      |
| 4                  |      |
| 5                  |      |
| 6                  |      |
| 7                  |      |
| 8                  |      |
| 9                  |      |
| 10                 |      |
| 11                 |      |
| <b>TOTAL</b>       |      |



Answer **all** questions in the spaces provided.

**0 1** This question is about the reactions of alkanes.

**0 1 . 1** Alkanes can be used as fuels.

Give an equation for the combustion of heptane ( $C_7H_{16}$ ) in an excess of oxygen.

[1 mark]



**0 1 . 2** Heptane can be obtained from the catalytic cracking of hexadecane ( $C_{16}H_{34}$ ) at a high temperature.

Identify a suitable catalyst for this process.

Give **one** condition other than high temperature.

Give an equation for the catalytic cracking of one molecule of hexadecane to produce one molecule of heptane, one molecule of cyclohexane and one other product.

[3 marks]

Catalyst Zeolite *→ very high surface area*

Condition Moderate Pressure



**0 1 . 3** Alkanes can be used in free-radical substitution reactions to produce halogenoalkanes.

Give equations for the propagation steps in the reaction of butane to form 2-chlorobutane.

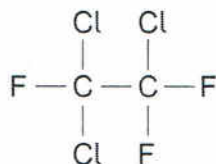
[2 marks]



0 1 4

Chlorofluorocarbons (CFCs) are a group of halogenoalkanes currently banned in many countries. They cannot be used as solvents or refrigerants because of their effect on the environment.

The structure of a CFC is shown.



Identify the radical produced from this CFC that is responsible for the depletion of ozone in the atmosphere.

Explain, with the aid of equations, why a single radical can cause the decomposition of many molecules of ozone.

[4 marks]

Radical



Explanation



The  $\text{Cl}^\bullet$  radical has been regenerated, and can be described as a catalyst.

10

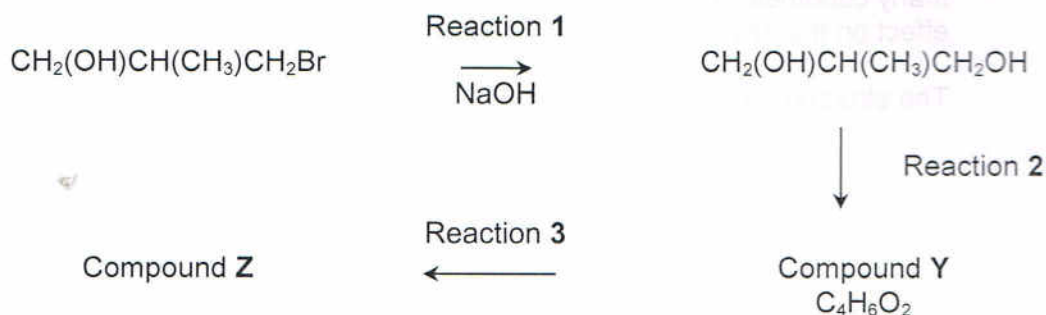
Turn over for the next question

Turn over ►



0 2

Halogenoalkanes are useful compounds in synthesis. A reaction pathway is shown.



0 2

1

Give the IUPAC name for  $\text{CH}_2(\text{OH})\text{CH}(\text{CH}_3)\text{CH}_2\text{Br}$

Bromo before methyl as alphabetical.

[1 mark]

3-bromo-2-methylpropan-1-ol

0 2

2

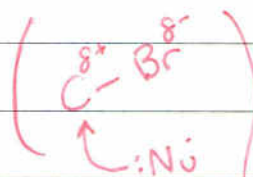
Reaction 1 occurs via a nucleophilic substitution mechanism.

Explain why the halogenoalkane is attacked by the nucleophile in this reaction.

[3 marks]

$\text{C}^{\delta+} - \text{Br}^{\delta-}$

Bromine is more electronegative than carbon which means that the C-Br bond is polar (the C is  $\delta^+$ ). This means the lone pair of electrons on the nucleophile can be donated to the  $\text{C}^{\delta+}$ .

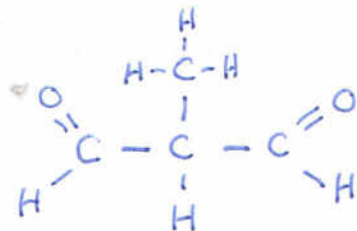




- 0 2 . 3 The infrared spectrum of Compound Y shows a significant absorption in the range  $1680-1750\text{ cm}^{-1}$

Draw the displayed formula of Compound Y.

[1 mark]



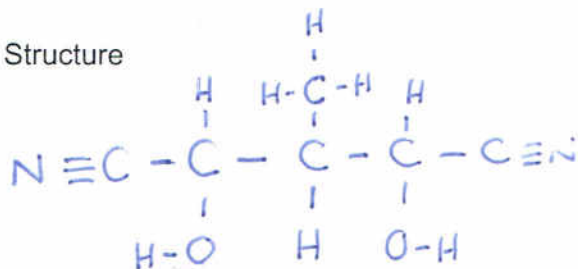
(two C=O gives strong peak @ 1680-1750)

- 0 2 . 4 Compound Z has the empirical formula  $\text{C}_3\text{H}_4\text{NO}$

- ① Give the structure of Compound Z.
- ② Suggest the reagent for Reaction 3.

[2 marks]

Structure



KCN and acid

↓  
gives  
:CN<sup>-</sup>  
nucleophile

Reagent for Reaction 3



0 3

The oxidation of propan-1-ol can form propanal and propanoic acid. The boiling points of these compounds are shown in **Table 1**.

**Table 1**

| Compound       | Boiling point / °C |
|----------------|--------------------|
| propan-1-ol    | 97                 |
| propanal       | 49                 |
| propanoic acid | 141                |

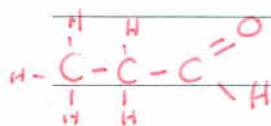
In a preparation of propanal, propan-1-ol is added dropwise to the oxidising agent and the aldehyde is separated from the reaction mixture by distillation.

0 3 . 1

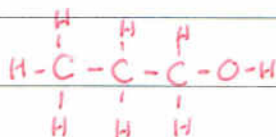
Explain, with reference to intermolecular forces, why distillation allows propanal to be separated from the other organic compounds in this reaction mixture.

**[3 marks]**

propanal which is an aldehyde would have dipole-dipole forces between molecules (as well as vdw). However both propan-1-ol and propanoic acid have Hydrogen bonding between molecules which are stronger than dipole-dipole in propanal, giving a lower boiling point and hence evaporated first.

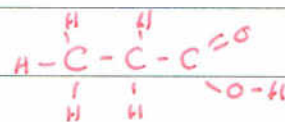


propanal

dipole only  
-dipole

propan-1-ol

H bonding



propanoic acid

H bonding



0 3 . 2 Give **two** ways of maximising the yield of propanal obtained by distillation of the reaction mixture. [2 marks]

1 Keep the temperature of the reaction mixture below the boiling point of propan-1-ol. This ensures that the propan-1-ol remains as a liquid.

2 Cool the distillate to ensure none evaporates away.

0 3 . 3 Describe how you would carry out a simple test-tube reaction to confirm that the sample of propanal obtained by distillation does **not** contain any propanoic acid. [2 marks]

could use Mg  
or any  
carbonate.

Add sodium carbonate to a sample of the distillate and if propanoic acid is present it will fizz (effervesce).

↓  
if carbonate  $\text{CO}_2$   
released.

Question 3 continues on the next page

Turn over ►



0 3 . 4

A student carried out an experiment to determine the enthalpy of combustion of ethanol.

Combustion of 457 mg of ethanol increased the temperature of 150 g of water from 25.1 °C to 40.2 °C

Calculate a value, in  $\text{kJ mol}^{-1}$ , for the enthalpy of combustion of ethanol in this experiment.

Give your answer to the appropriate number of significant figures.

(The specific heat capacity of water is  $4.18 \text{ J K}^{-1} \text{ g}^{-1}$ )

[3 marks]

$$q = m \times c \times \Delta T$$

← mass of water

$$q = 150 \times 4.18 \times 15.1$$

$$q = 9467.7 \text{ J} \quad \div 1000 = 9.4677 \text{ kJ}$$

$$\frac{457 \text{ mg}}{1000} = 0.457 \text{ g}$$

$$\frac{0.457}{46} = 9.93 \times 10^{-3} \text{ mole}$$

↑  
Mr of ethanol

$$\frac{9.4677}{9.93 \times 10^{-3}} = \underline{\underline{-953.4 \text{ kJ mol}^{-1}}} \quad \therefore \underline{\underline{-953 \text{ kJ mol}^{-1}}}$$

3 S.F.

Enthalpy of combustion \_\_\_\_\_  $\text{kJ mol}^{-1}$





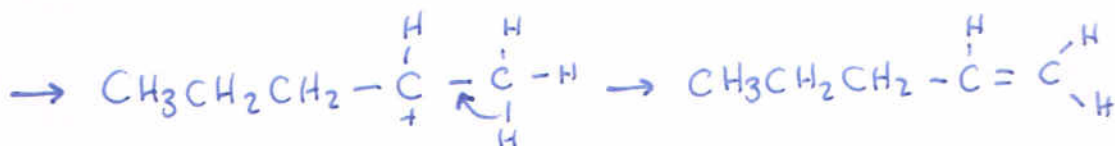
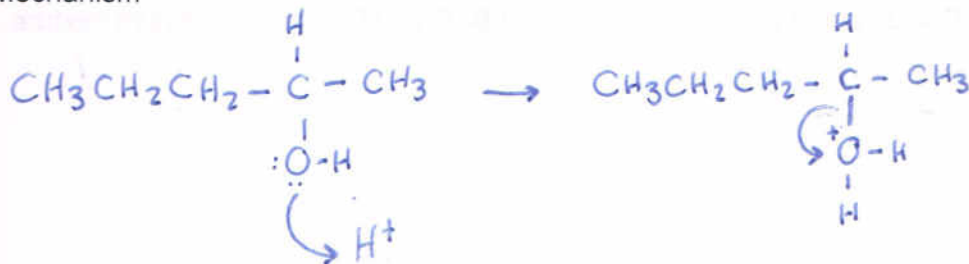
- 03.5 A mixture of isomeric alkenes is produced when pentan-2-ol is dehydrated in the presence of hot concentrated sulfuric acid. Pent-1-ene is one of the isomers produced.

Name and outline a mechanism for the reaction producing pent-1-ene.

[4 marks]

Name of mechanism Elimination

Mechanism



Make sure you  
form pent-1-ene  
not pent-2-ene!

- 03.6 A pair of stereoisomers is also formed in the reaction in Question 03.5.

Name the less polar stereoisomer formed.

Explain how this type of stereoisomerism arises.

not seen this question before!

2 isomers usually more polar as more unevenly distributed.

[2 marks]

Name E-pent-2-ene



Explanation restricted rotation around C=C

bond, and each carbon in double bond has  
two different groups attached.



0 4

Compounds **A** and **B** react together to form an equilibrium mixture containing compounds **C** and **D** according to the equation



0 4 . 1

A beaker contained  $40 \text{ cm}^3$  of a  $0.16 \text{ mol dm}^{-3}$  aqueous solution of **A**.  $9.5 \times 10^{-3} \text{ mol}$  of **B** and  $2.8 \times 10^{-2} \text{ mol}$  of **C** were added to the beaker and the mixture was left to reach equilibrium.

The equilibrium mixture formed contained  $3.9 \times 10^{-3} \text{ mol}$  of **A**.

Calculate the amounts, in moles, of **B**, **C** and **D** in the equilibrium mixture.

[5 marks]

|                                 | $2A + B \rightleftharpoons 3C + D$ |  |   |                       |
|---------------------------------|------------------------------------|--|---|-----------------------|
| I<br>moles                      | $6.4 \times 10^{-3}$               | $9.5 \times 10^{-3}$   | $2.8 \times 10^{-2}$  | 0                     |
| Eq                              | $3.9 \times 10^{-3}$               | $9.5 \times 10^{-3} - (1.25 \times 10^{-3}) = 8.25 \times 10^{-3}$ | $2.8 \times 10^{-2} + (1.25 \times 10^{-3} \times 3) = 0.03175$ | $1.25 \times 10^{-3}$ |
| Conc ( $\div \frac{40}{1000}$ ) | 0.0975                             | 0.18875  | 0.13375   | 0.18875               |

$$C = \frac{m}{V} \quad C \times V = m \quad 0.16 \times \left( \frac{40}{1000} \right) = 6.4 \times 10^{-3} \text{ moles of A.}$$

The key to this is to work out an  $x$ .  $6.4 \times 10^{-3} - 3.9 \times 10^{-3} = 2.5 \times 10^{-3} = 2x$   
 $x = \frac{2.5 \times 10^{-3}}{2}$

Then use this to work out the rest.  $\leftarrow = 1.25 \times 10^{-3}$

Amount of **B** \_\_\_\_\_ mol

Amount of **C** \_\_\_\_\_ mol

Amount of **D** \_\_\_\_\_ mol

0 4 . 2

Give the expression for the equilibrium constant ( $K_c$ ) for this equilibrium and its units.

[2 marks]

$$K_c = \frac{[C]^3 \times [D]}{[A]^2 \times [B]}$$

$$\frac{\text{mol dm}^{-3} \times \text{mol dm}^{-3} \times \text{mol dm}^{-3} \times \text{mol dm}^{-3}}{\text{mol dm}^{-3} \times \text{mol dm}^{-3} \times \text{mol dm}^{-3}}$$

Units  $\text{mol dm}^{-3}$



0 4 . 3

A different equilibrium mixture of these four compounds, at a different temperature, contained 0.21 mol of **B**, 1.05 mol of **C** and 0.076 mol of **D** in a total volume of  $5.00 \times 10^2 \text{ cm}^3$  of solution.

At this temperature the numerical value of  $K_c$  was 116

Calculate the concentration of **A**, in  $\text{mol dm}^{-3}$ , in this equilibrium mixture.  
Give your answer to the appropriate number of significant figures.

[3 marks]

$$K_c = \frac{[C]^3 \times [D]}{[A]^2 \times [B]}$$

$$116 = \frac{(1.05)^3 \times \left(\frac{0.076}{0.5}\right)}{[A]^2 \times \left(\frac{0.21}{0.5}\right)}$$

Don't forget  
to divide by  
volume!!

$$116 = \frac{9.261 \times 0.152}{[A]^2 \times 0.42}$$

$$\sqrt{\frac{9.261 \times 0.152}{116 \times 0.42}} = [A]$$

$$[A] = 0.1699 \text{ or } \underline{\underline{0.17}} \text{ 2sf}$$

lowest no.  
of sig figs in  
question.

Concentration of **A** \_\_\_\_\_  $\text{mol dm}^{-3}$

0 4 . 4

Justify the statement that adding more water to the equilibrium mixture in Question 04.3 will lower the amount of **A** in the mixture.

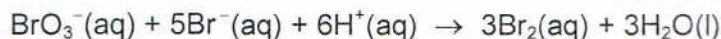
[3 marks]

Adding more water will cause all concentrations to decrease, however equilibrium will shift to the side with more moles of substance which is right and hence lowering the amount of **A**.





0 5 Bromate(V) ions and bromide ions react in acid conditions according to the equation



0 5 . 1 A series of experiments was carried out at a given temperature. The results were used to deduce the rate equation for the reaction.

$$\text{rate} = k [\text{BrO}_3^-][\text{Br}^-][\text{H}^+]^2$$

Table 2 shows an incomplete set of results.

Table 2

| Experiment | Initial $[\text{BrO}_3^-]$<br>/ $\text{mol dm}^{-3}$ | Initial $[\text{Br}^-]$<br>/ $\text{mol dm}^{-3}$ | Initial $[\text{H}^+]$<br>/ $\text{mol dm}^{-3}$ | Initial rate of reaction<br>/ $\text{mol dm}^{-3} \text{s}^{-1}$ |
|------------|--|---|--|--|
| 1          | 0.10   | 0.20  | 0.30   | $2.4 \times 10^{-2}$   |
| 2          | 0.15   | 0.20  | 0.30   | $3.6 \times 10^{-2}$   |
| 3          | 0.20   | 0.40  | 0.50   | 0.27   |
| 4          | 0.10   | 0.10  | 0.45   | $2.7 \times 10^{-2}$   |

Use the data from Experiment 1 to calculate a value for the rate constant,  $k$ , at this temperature and give its units.

Give your answer to an appropriate number of significant figures.

[3 marks]

$$2.4 \times 10^{-2} = k \times 0.1 \times 0.20 \times (0.30)^2$$

$$\frac{2.4 \times 10^{-2}}{0.1 \times 0.2 \times 0.3^2} = 13.33$$

$$\frac{\text{mol dm}^{-3} \text{s}^{-1}}{\text{mol dm}^{-3} \times \text{mol dm}^{-3} \times \text{mol dm}^{-3}} = \text{mol}^3 \text{dm}^9 \text{s}^{-1}$$

$k$  \_\_\_\_\_

Units \_\_\_\_\_

0 5 . 2 Complete Table 2.

[3 marks]

Space for working

$$\sqrt{\frac{2.7 \times 10^{-2}}{13.33 \times 0.10 \times 0.10}} = [\text{H}^+] = 0.45$$

Question 5 continues on the next page

Turn over ►





0 5 . 3

A second series of experiments was carried out to investigate how the rate of the reaction varies with temperature. The results were used to obtain a value for the activation energy of the reaction,  $E_a$ .

Identical amounts of reagents were mixed at different temperatures. The time taken,  $t$ , for a fixed amount of bromine to be formed was measured at different temperatures.

The results are shown in **Table 3**.

**Table 3**

| Temperature, $T$<br>/ K | $\frac{1}{T} / \text{K}^{-1}$ | Time, $t$<br>/ s | $\frac{1}{t} / \text{s}^{-1}$ | $\ln \frac{1}{t}$ |
|-------------------------|-------------------------------|------------------|-------------------------------|-------------------|
| 286                     | $3.50 \times 10^{-3}$         | 54               | $1.85 \times 10^{-2}$         | -3.99             |
| 295                     | $3.39 \times 10^{-3}$         | 27               | $3.70 \times 10^{-2}$         | -3.30             |
| 302                     | $3.31 \times 10^{-3}$         | 15               | $6.67 \times 10^{-2}$         | -2.71             |
| 312                     | $3.21 \times 10^{-3}$         | 8                | $1.25 \times 10^{-1}$         | -2.08             |

Complete **Table 3**.

[2 marks]

0 5 . 4

The Arrhenius equation can be written as

$$\ln k = -\frac{E_a}{R} \left( \frac{1}{T} \right) + C_1$$

In this experiment, the rate constant,  $k$ , is directly proportional to  $\frac{1}{t}$

Therefore

$$\ln \frac{1}{t} = -\frac{E_a}{R} \left( \frac{1}{T} \right) + C_2$$

where  $C_1$  and  $C_2$  are constants.

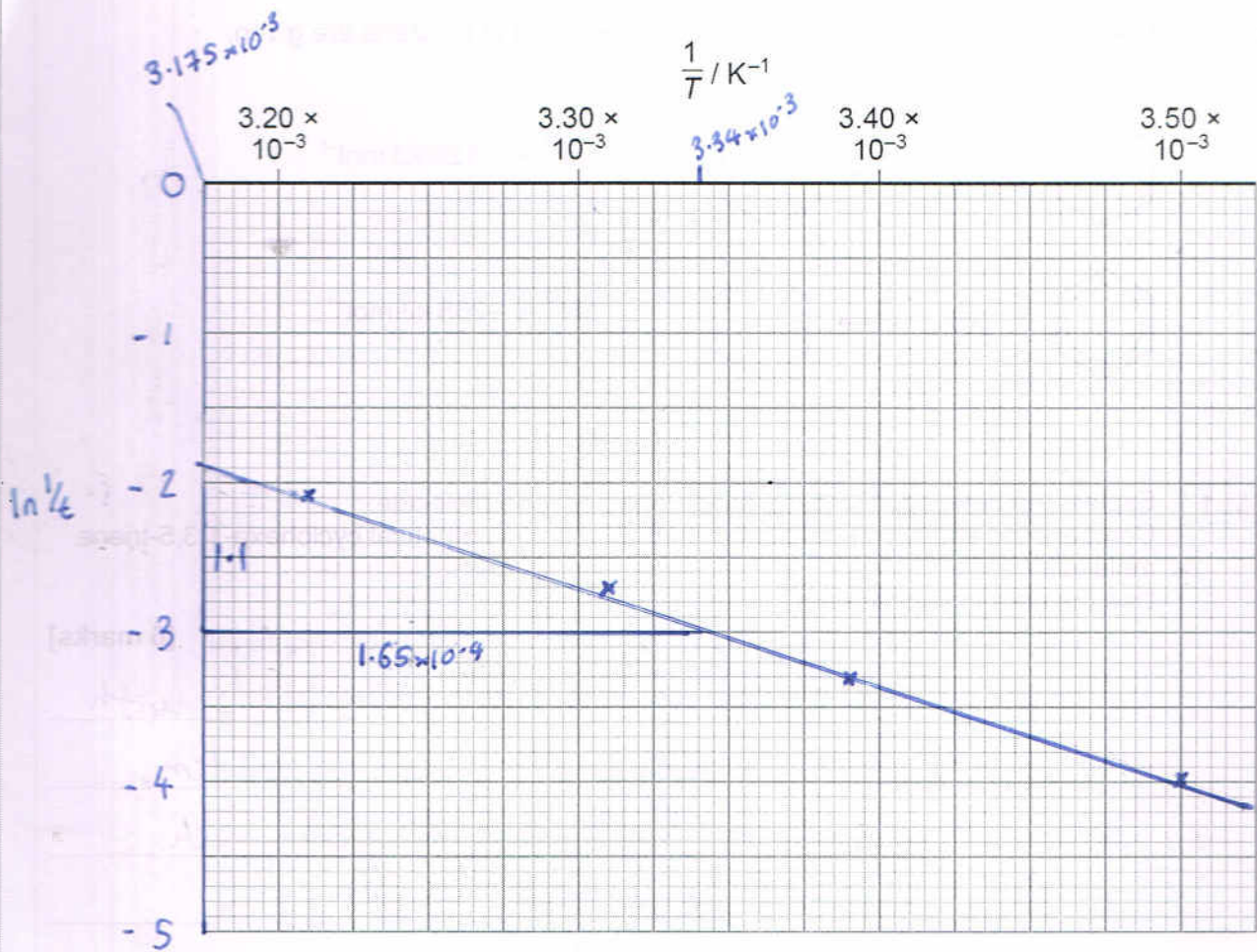
Use values from **Table 3** to plot a graph of  $\ln \frac{1}{t}$  (y axis) against  $\frac{1}{T}$  on the grid.

Use your graph to calculate a value for the activation energy, in  $\text{kJ mol}^{-1}$ , for this reaction.

The value of the gas constant,  $R = 8.31 \text{ J K}^{-1} \text{ mol}^{-1}$

[6 marks]





$$\text{gradient} = \frac{-1.1}{1.65 \times 10^{-4}} = -6.66 \times 10^3 \text{ or } \underline{\underline{-6667}}$$

$$E_a = 6667 \times 8.31 = 55400 \text{ J mol}^{-1}$$

$$\therefore \frac{55400}{1000} = \underline{\underline{55.4 \text{ kJ mol}^{-1}}}$$

\* gradient of this plot is equal to  $\frac{-E_a}{R}$

\* if asked, y intercept is equal to  $\ln A$

Activation energy

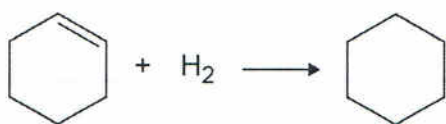
$\text{kJ mol}^{-1}$

14

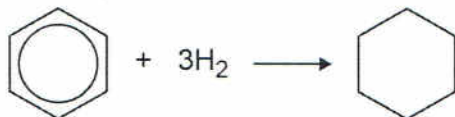


0 6

Data about the hydrogenation of cyclohexene and of benzene are given.



$$\Delta H^\ominus = (-120 \text{ kJ mol}^{-1} \times 3)$$



$$\Delta H^\ominus = -208 \text{ kJ mol}^{-1}$$

↓ less exothermic

0 6

1

① Explain the bonding in and the shape of a benzene molecule.

② Compare the stability of benzene with that of the hypothetical cyclohexa-1,3,5-triene molecule.

Use the data in your answer.

[6 marks]

- ① In benzene each carbon atom has three bonds with the spare electrons forming a delocalised  $\pi$  system.
- ② Benzene is planar in shape and hexagonal. each carbon-carbon bond is of equal length.
- ③ Benzene is more stable than cyclohexa-1,3,5-triene as shown by the  $\Delta H^\ominus$  values of  $-208 \text{ kJ mol}^{-1}$  for benzene and  $3 \times -120 = -360 \text{ kJ mol}^{-1}$  for cyclohexa-1,3,5-triene. These values suggest that if ring system for benzene was broken and replaced with new C-C bonds it is less exothermic than expected and hence more stable.





0 6 . 2

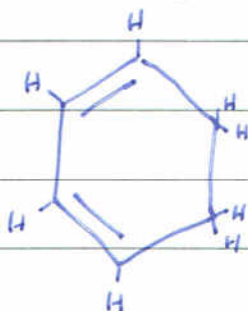
The enthalpy of hydrogenation of cyclohexa-1,3-diene is **not** exactly double that of cyclohexene.

Suggest a value for the enthalpy of hydrogenation of cyclohexa-1,3-diene and justify your value.

[3 marks]

*(-239 - 121)* →  $-180 \text{ kJ mol}^{-1}$  is a good estimate. This is because the diene has delocalisation itself because the two double bonds are conjugated (they have a single bond separating two double bonds).

*Benzene has this delocalised stability even more.*





07

Acyl chlorides are useful reagents in synthesis. They react with aromatic compounds and also with alcohols.

07.1

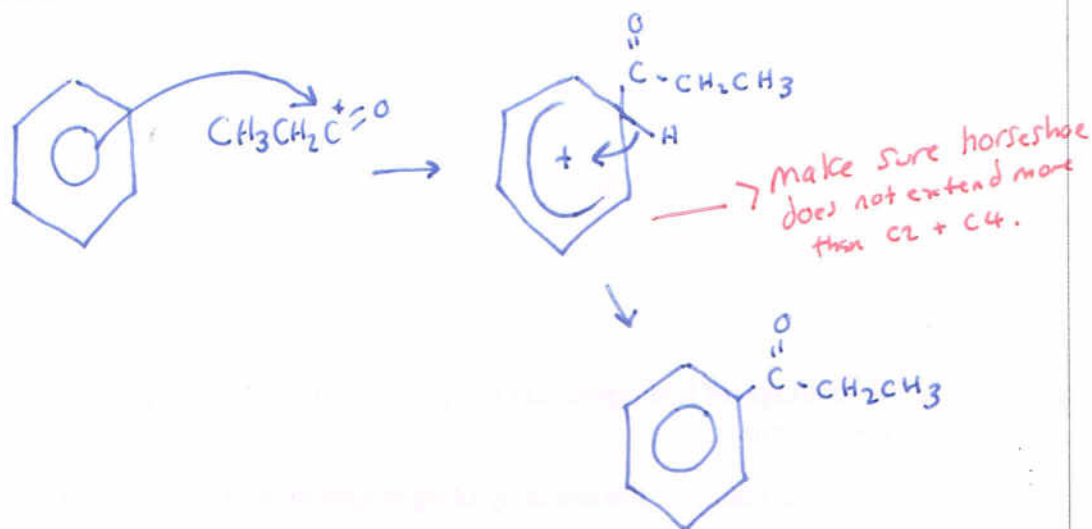
$\text{CH}_3\text{CH}_2\text{COCl}$  reacts with benzene in the presence of  $\text{AlCl}_3$  in an electrophilic substitution reaction.

Give an equation for the reaction of  $\text{CH}_3\text{CH}_2\text{COCl}$  with  $\text{AlCl}_3$  to form the electrophile. Outline a mechanism for the reaction of this electrophile with benzene.

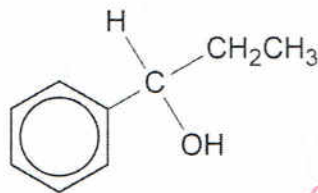
[4 marks]



Mechanism



07.2 The organic product in Question 07.1 can be converted into the alcohol shown.



*this is usually asked for non-aromatic compounds but some applies*

Give the IUPAC name of the alcohol.

Give the reagent needed for this reaction and name the mechanism.

[3 marks]

IUPAC name 1-phenylpropan-1-ol

Reagent NaBH<sub>4</sub>

Name of mechanism Nucleophilic addition

07.3 The alcohol shown in Question 07.2 reacts with ethanoyl chloride to form an ester.

Describe what would be observed when the alcohol reacts with ethanoyl chloride.

Name the mechanism for the reaction to form the ester.

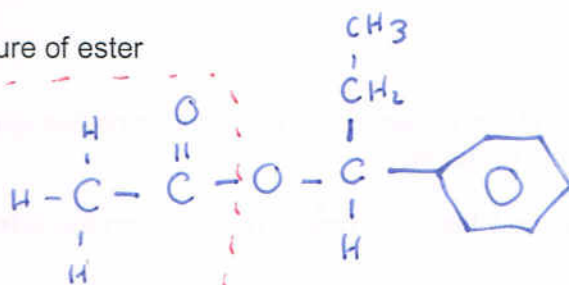
Draw the structure of the ester.

Observation Steamy fumes

*HCl gas released.* [3 marks]

Name of mechanism Addition-elimination

Structure of ester



*ethanoyl chloride.*

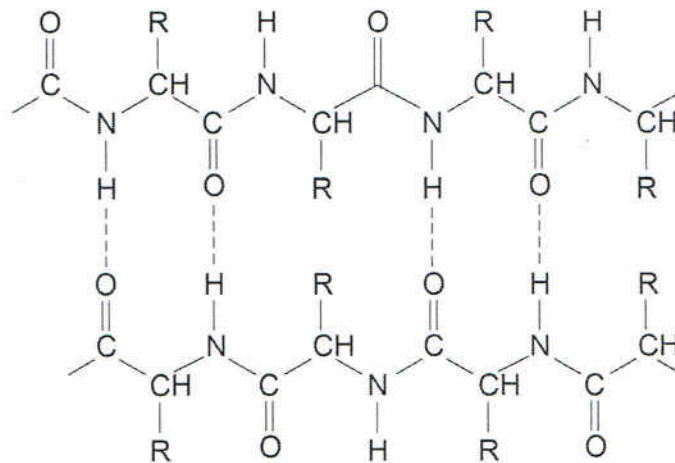
Turn over for the next question



0 8

Use the Data Booklet to help you answer this question about amino acids.  
Figure 1 shows parts of two polypeptide chains in a beta-pleated sheet of a protein.

Figure 1



0 8 . 1

The polypeptide chains are held together by hydrogen bonding as shown in Figure 1.

Explain how these hydrogen bonds form.

[2 marks]

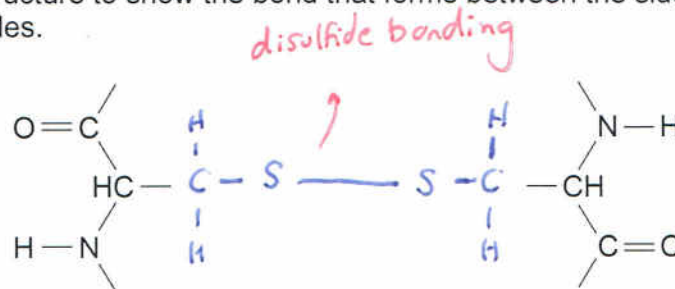
H atoms are  $\delta^+$  and are attracted to lone pair of electrons on oxygen atoms.

0 8 . 2

A different type of bond can form between two polypeptide chains when the chains each contain the amino acid cysteine.

Complete the structure to show the bond that forms between the side chains of two cysteine molecules.

[1 mark]



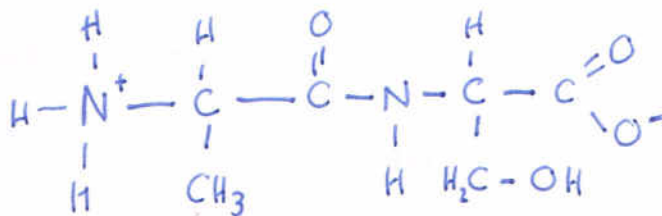
- 0 8 . 3 The type of bond in Question 08.2 between two polypeptide chains influences the three-dimensional structure of the protein.

Name this type of protein structure.

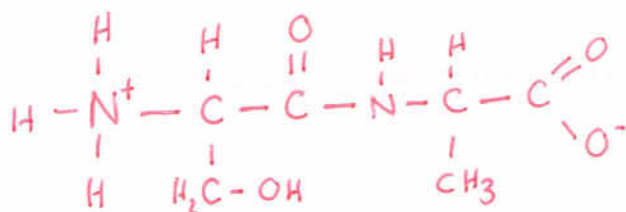
[1 mark]

Tertiary (or Quaternary)

- 0 8 . 4 Draw the structure of the zwitterion of a dipeptide formed by alanine and serine. [2 marks]



OR



Turn over for the next question





0 9

Use the Data Booklet to help you answer this question about DNA.

**Figure 2** shows a fragment of a DNA double helix.

The letters A, C, G and T represent the four bases in one strand.

The numbers 1, 2, 3, 4 and 5 represent the bases in the complementary strand.

**Figure 2**

|   |   |   |   |   |
|---|---|---|---|---|
| A | C | G | T | C |
| 1 | 2 | 3 | 4 | 5 |

A-T  
C-G

0 9 . 1

Complete **Table 4** to show the correct sequence of bases in the complementary strand represented by the numbers 1 to 5

[1 mark]

**Table 4**

|   |   |   |   |   |
|---|---|---|---|---|
| 1 | 2 | 3 | 4 | 5 |
| T | G | C | A | G |

0 9 . 2

Deduce the total number of hydrogen bonds formed between the five bases in each strand.

Tick (✓) **one** box.

[1 mark]

|    |    |    |    |
|----|----|----|----|
| 10 | 12 | 13 | 15 |
|    |    | ✓  |    |

A-T = 2 hydrogen bonds

C-G = 3 hydrogen bonds

∴ two A-T = 4

∴ three C-G = 9

= 13



0 9 . 3

Base A is part of a nucleotide in the DNA strand shown in **Figure 2**.

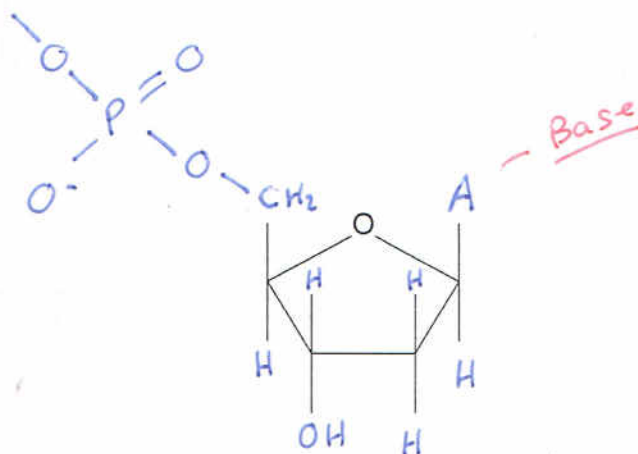
A nucleotide contains a 2-deoxyribose molecule.

An incomplete 2-deoxyribose molecule is shown.

Complete the structure to show the nucleotide that contains base A.

You should represent base A by the letter A.

[2 marks]



4

Turn over for the next question

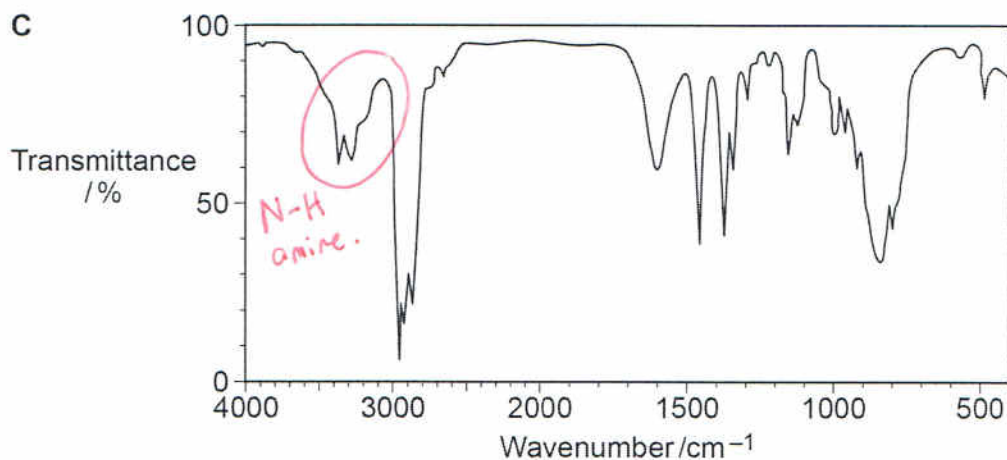
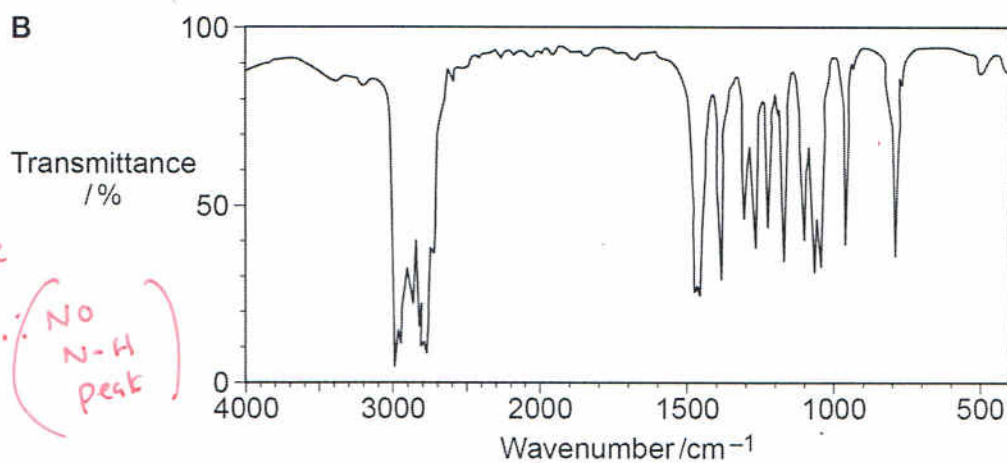
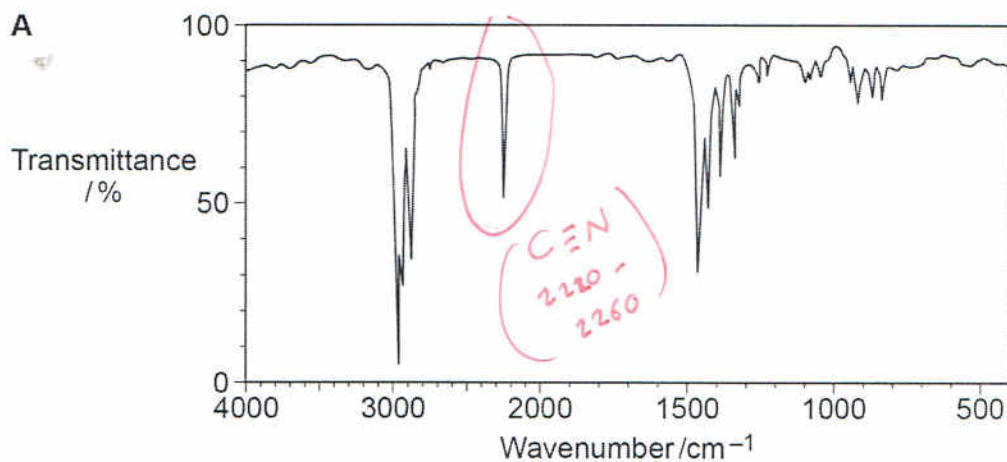
Turn over ►



1 0

This question is about amines.

1 0 . 1

The infrared spectra **A**, **B** and **C** are those of a primary amine, a tertiary amine and a nitrile, but not necessarily in that order.

Give the letter of each compound in the correct box.

[1 mark]

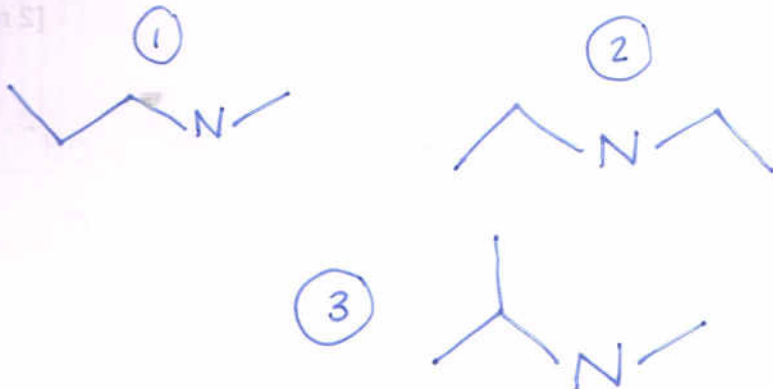
| primary amine | tertiary amine | nitrile |
|---------------|----------------|---------|
| C             | B              | A       |



1 0 . 2 There are **three** secondary amines that contain four carbon atoms per molecule.

Draw the skeletal formulas of these **three secondary amines**.

[2 marks]



1 0 . 3 Primary amines can be prepared by the reaction of halogenoalkanes with ammonia or by the reduction of nitriles.

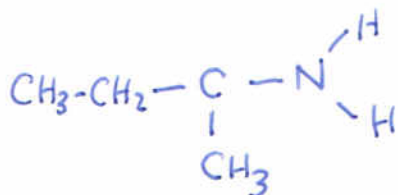
Justify the statement that it is better to prepare primary amines from nitriles rather than from halogenoalkanes.

[2 marks]

If you use haloalkanes with ammonia you are likely to get multiple substitutions forming N-substituted amines. Whereas with nitriles you do not.

1 0 . 4 Draw the structure of a primary amine with four carbon atoms that **cannot** be formed from a nitrile.

[1 mark]



any 2° or 3° amine. nitriles are primary only.

Turn over ▶





1 0 . 5 A student dissolves a few drops of propylamine in  $1 \text{ cm}^3$  of water in a test tube.

Give an equation for the reaction that occurs.

Describe what is observed when Universal Indicator is added to this solution.

[2 marks]

(amines are basic)

Equation  $\text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2 + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_3^+ + \text{OH}^-$

Observation universal indicator will turn from green to blue/purple.

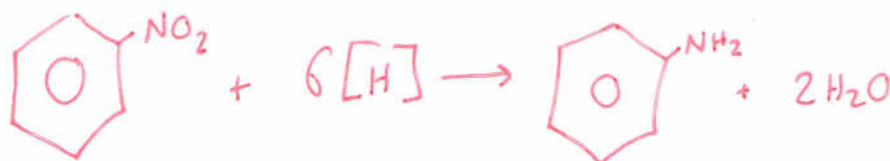
1 0 . 6 Phenylamine can be prepared by a process involving the reduction of nitrobenzene using tin and an excess of hydrochloric acid.

- ① Give an equation for the reduction of nitrobenzene to form phenylamine. Use [H] to represent the reducing agent.
- ② Explain why an aqueous solution is obtained in this reduction even though phenylamine is insoluble in water.

[2 marks]

Equation  $\text{C}_6\text{H}_5\text{NO}_2 + 6[\text{H}] \rightarrow \text{C}_6\text{H}_5\text{NH}_2 + 2\text{H}_2\text{O}$

Explanation This is because an ionic salt is formed,  $\text{C}_6\text{H}_5\text{NH}_3^+\text{Cl}^-$ .

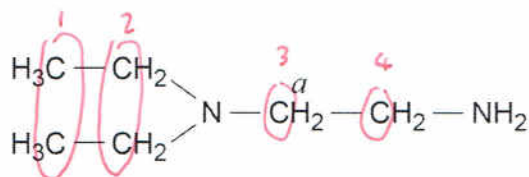


1 1

There are several isomers with the molecular formula  $C_6H_{16}N_2$ 

1 1 . 1

One isomer is shown.

Give the number of peaks in the  $^{13}C$  NMR spectrum of this isomer.State and explain the splitting pattern of the peak for the hydrogens labelled *a* in its  $^1H$  NMR spectrum.*(careful here - starts as  $^{13}C$  NMR then  $^1H$  NMR!)*

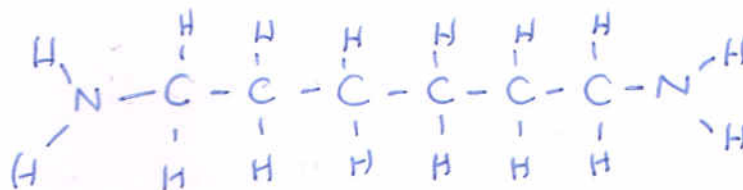
[3 marks]

Number of  $^{13}C$  peaks 4 peaksSplitting pattern  $\alpha$  = triplet.Explanation This is due to the two hydrogens next to it (adjacent), and following the  $n+1$  rule.

1 1 . 2

Draw the structure of the isomer of  $C_6H_{16}N_2$  used to make nylon 6,6

[1 mark]

1,6 diamino hexane

Question 11 continues on the next page

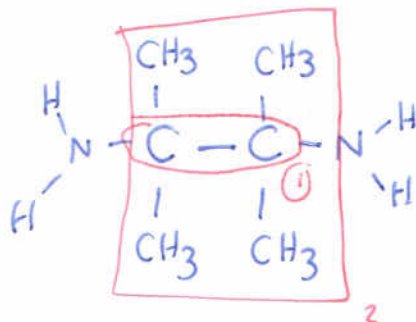
Turn over ►



1 1 . 3

Draw the structure of the isomer of  $C_6H_{16}N_2$  that contains two **primary** amine groups and has only two peaks in its  $^{13}C$  NMR spectrum.

[1 mark]

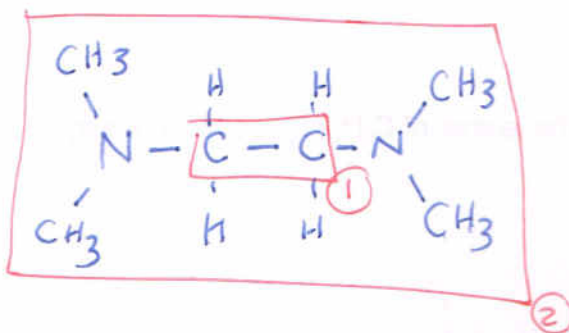
② CH<sub>3</sub>'s are equivalent

① C's are equivalent

1 1 . 4

Draw the structure of the isomer of  $C_6H_{16}N_2$  that contains two **tertiary** amine groups and has only two peaks in its  $^{13}C$  NMR spectrum.

[1 mark]



① C's equivalent

② CH<sub>3</sub>'s are equivalent.

END OF QUESTIONS

6

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