



3. The rate constants for a reaction at a variety of temperatures are shown in the table below. Use the data to create an Arrhenius plot and therefore work out the activation energy and the Arrhenius constant.

<b>T / K</b>	<b>k / s<sup>-1</sup></b>		
<b>150</b>	4.59E-48		
<b>200</b>	2.44E-36		
<b>250</b>	5.90E-29		
<b>300</b>	9.60E-24		
<b>350</b>	1.05E-20		
<b>400</b>	4.25E-18		