

C2 Exothermic and Endothermic Quiz Answers

1. How can you show that a reaction is exothermic?

Use a thermometer to find the temperature of the surroundings before and after the reaction. If the temperature rises, the reaction is exothermic. If the temperature drops, the reaction is endothermic.

2. What is the meaning of endothermic?

Energy is taken from the surroundings to the chemical system.

3. In an endothermic reaction the temperature of the surroundings drops. Why?

Energy is transferred from the surroundings to the chemical system.

4. Give examples of endo and exothermic reactions.

Combustion and respiration are exothermic. Thermal decomposition and photosynthesis are endothermic.

5. If a forward reaction is exothermic, what do you need to do to reverse the reaction?

Add heat as the reverse reaction will be endothermic.

6. How do you know from an energy profile diagram that a reaction is exothermic?

The energy of the products is below the energy of the reactants.

7. How do you know from an energy profile diagram that a reaction is endothermic?

The energy of the products is above the energy of the reactants.

8. How do you convert blue hydrated copper sulphate into white anhydrous copper sulphate?

Supply heat to evaporate the water present in the blue hydrated copper sulfate.

9. Give a use for anhydrous copper sulphate.

To detect/absorb moisture. It will turn blue if water is present.