

C2 Acids, Bases and Salts Quiz

1. Suggest a pH value for hydrochloric acid.
2. What ions make ethanoic acid acidic?
3. What ions make ammonia solution alkaline?
4. Which acid is needed to make ammonium nitrate?
5. What is the formula of ammonium sulphate?
6. Suggest a pH value for ammonia solution.
7. What type of reaction occurs between sulphuric acid and ammonia?
8. Why do farmers use ammonium salts on their fields?
9. What is a precipitate?
10. What do plants use the nitrogen in fertilisers for?
11. How can solid lead iodide be separated from solution?
12. How can copper sulphate crystals be separated from copper sulfate solution?

13. Why is KOH a strong alkali?
14. Why is ethanoic acid a weak acid?
15. What is the Brønsted- Lowry definition of an acid and base?
16. Write an ionic neutralisation equation including state symbols.
17. What are the four state symbols and what do they mean?
18. What is produced when an acid reacts with a metal oxide?
19. What is produced when an acid reacts with a metal hydroxide?
20. What is produced when an acid reacts with a metal?
21. What is produced when an acid reacts with a metal carbonate or metal hydrogen carbonate?
22. What is produced when an acid reacts with ammonia?
23. What is the difference between ammonia and ammonium?
24. When copper sulphate is made by reacting copper oxide with sulphuric acid, the acid is heated. Why?
25. How would you remove unreacted copper oxide from solution?

26. Name the salt formed from hydrochloric acid.
27. Name the salt formed from sulphuric acid.
28. Name the salt formed from nitric acid.
29. Why is dry hydrogen chloride gas not acidic?
30. Why is NaCl neutral?
31. How do you make a soluble salt from an acid and an alkali?
32. How do you make a soluble salt from an acid and a solid base?
33. Which salts are insoluble?
34. Which salts are soluble?
35. How does UI show the difference in acid strength when added to ethanoic acid and hydrochloric acid of same concentration?