

Write your name here

Surname

Other names

Pearson
Edexcel GCSE

Centre Number

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Candidate Number

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Chemistry/Science

Unit C1: Chemistry in Our World

Higher Tier

Thursday 19 May 2016 – Morning

Time: 1 hour

Paper Reference

5CH1H/01

You must have:

Calculator, ruler

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*

Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*
- Questions labelled with an **asterisk** (*) are ones where the quality of your written communication will be assessed
– *you should take particular care with your spelling, punctuation and grammar, as well as the clarity of expression, on these questions.*

Advice

- Read each question carefully before you start to answer it.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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The Periodic Table of the Elements

1	2	3	4	5	6	7	0										
7 Li lithium 3	9 Be beryllium 4	11 Na sodium 11	12 C carbon 6	13 Al aluminium 13	14 N nitrogen 7	15 P phosphorus 15	16 O oxygen 8	17 F fluorine 9	18 Ne neon 10								
19 K potassium 19	20 Ca calcium 20	23 Sc scandium 21	24 Ti titanium 22	25 V vanadium 23	26 Cr chromium 24	27 Mn manganese 25	28 Fe iron 26	29 Co cobalt 27	30 Ni nickel 28	31 Cu copper 29	32 Zn zinc 30	33 Ga gallium 31	34 Ge germanium 32	35 As arsenic 33	36 Se selenium 34	37 Br bromine 35	38 Kr krypton 36
37 Rb rubidium 37	38 Sr strontium 38	39 Y yttrium 39	40 Zr zirconium 40	41 Nb niobium 41	42 Mo molybdenum 42	43 Tc technetium 43	44 Ru ruthenium 44	45 Rh rhodium 45	46 Pd palladium 46	47 Ag silver 47	48 Cd cadmium 48	49 In indium 49	50 Sn tin 50	51 Sb antimony 51	52 Te tellurium 52	53 I iodine 53	54 Xe xenon 54
55 Cs caesium 55	56 Ba barium 56	57 La* lanthanum 57	72 Hf hafnium 72	73 Ta tantalum 73	74 W tungsten 74	75 Re rhenium 75	76 Os osmium 76	77 Ir iridium 77	78 Pt platinum 78	79 Au gold 79	80 Hg mercury 80	81 Tl thallium 81	82 Pb lead 82	83 Bi bismuth 83	84 Po polonium 84	85 At astatine 85	86 Rn radon 86
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111	Elements with atomic numbers 112-116 have been reported but not fully authenticated						

1	H hydrogen 1
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relative atomic mass
atomic symbol
name
atomic (proton) number

* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.

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Questions begin on next page.



Answer ALL questions

Some questions must be answered with a cross in a box ☒. If you change your mind about an answer, put a line through the box ~~☒~~ and then mark your new answer with a cross ☒.

Carbon dioxide

- 1 (a) The Earth's early atmosphere contained larger amounts of water vapour and carbon dioxide than it does today.

Explain how the amounts of water vapour and carbon dioxide in the Earth's atmosphere decreased.

(2)

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- (b) Describe how the presence of carbon dioxide in the atmosphere helps to keep the Earth warm.

(1)

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- (c) State a human activity, apart from burning fuels, that has increased the percentage of carbon dioxide in the atmosphere.

(1)

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(d) This table shows a set of data for the percentage of carbon dioxide in the atmosphere and the mean surface temperature of the Earth in the years 1960 and 2014.

year	percentage of carbon dioxide	mean surface temperature / °C
1960	0.0318	14.0
2014	0.0401	14.4

Suggest why this information does not prove that the increase in percentage of carbon dioxide alone has caused the increase in the mean surface temperature of the Earth.

(1)

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(e) Hydrogen can be used as a fuel for cars.

(i) State one advantage and one disadvantage of using a car which uses hydrogen as the fuel rather than a car which uses petrol as the fuel.

(2)

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(ii) Write the word equation for the reaction between hydrogen and oxygen.

(1)

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(Total for Question 1 = 8 marks)



Rocks

- 2 (a) Granite is an igneous rock.

The size of crystals in granite vary.

Describe how a sample of granite which contains large crystals was formed.

(2)

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- (b) Limestone and marble are natural forms of calcium carbonate.

Limestone is a sedimentary rock.

Marble is a metamorphic rock, formed from limestone.

Describe how marble is formed from limestone.

(2)

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- (c) Calcium carbonate is used to treat waste gases in coal-fired power stations.

Explain why calcium carbonate can be used in this way.

(2)

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- (d) (i) When calcium carbonate is heated, it breaks down to form calcium oxide and carbon dioxide.

What type of reaction is this?

Put a cross (☒) in the box next to your answer.

(1)

- A combustion
- B decomposition
- C oxidation
- D precipitation

- (ii) Calcium oxide reacts with water to form calcium hydroxide, Ca(OH)_2 .

Write the balanced equation for the reaction between calcium oxide and water.

(2)

(Total for Question 2 = 9 marks)



Acids and electrolysis

3 (a) Which of these substances is produced in the stomach to help digestion?

Put a cross (☒) in the box next to your answer.

(1)

- A calcium oxide
- B hydrochloric acid
- C magnesium oxide
- D sulfuric acid

(b) Nitric acid reacts with magnesium carbonate to form a salt, water and a gas.

(i) State the name of the salt formed in this reaction.

(1)

(ii) Which of these is the gas produced in this reaction?

Put a cross (☒) in the box next to your answer.

(1)

- A carbon dioxide
- B hydrogen
- C oxygen
- D nitrogen

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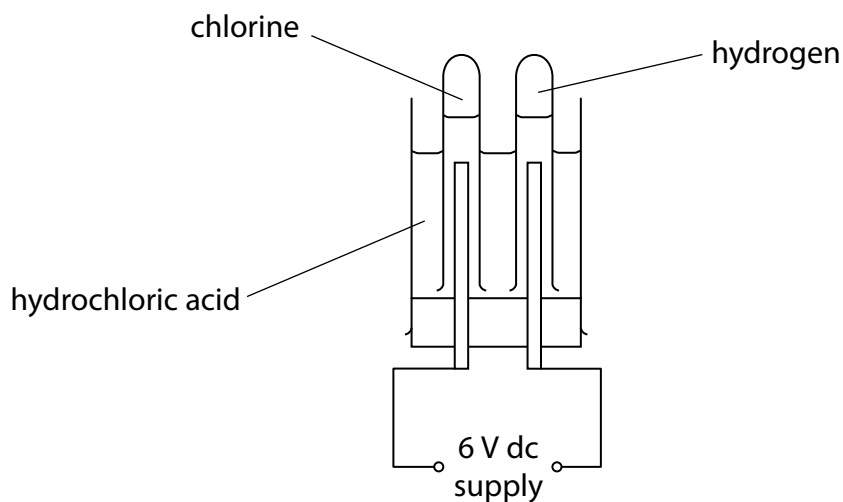


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(c) Hydrochloric acid is electrolysed using this apparatus.
Hydrogen and chlorine are produced.



(i) Describe the test to show that a gas is chlorine.

(2)

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(ii) Chlorine is a toxic gas.

State a safety precaution that should be taken when chlorine gas is formed in a reaction.

(1)

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(iii) Write the balanced equation for the decomposition of hydrochloric acid to form hydrogen and chlorine.

(3)

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(Total for Question 3 = 9 marks)



Alkenes and polymers

4 (a) Complete the sentence by putting a cross (☒) in the box next to your answer. (1)

Alkenes are

- A saturated carbohydrates
- B saturated hydrocarbons
- C unsaturated carbohydrates
- D unsaturated hydrocarbons

(b) Propene is an alkene and has the molecular formula C_3H_6 .

(i) Draw the structure of a propene molecule, showing all of the bonds. (2)

(ii) Propane is an alkane.
Propane and propene are both gases.

Given a sample of each gas, describe a test, carried out on both gases, to show which gas is propane and which gas is propene.

(3)

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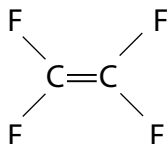
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(c) The diagram shows the structure of a tetrafluoroethene molecule.



Tetrafluoroethene can form the polymer PTFE.

(i) Draw a diagram to show part of a PTFE molecule formed from two tetrafluoroethene molecules.

(2)

(ii) PTFE does not allow other substances to stick to it.
State a use of PTFE related to this property.

(1)

(d) Many polymers cause problems because they persist for a long time when they are put in landfill sites.

State an alternative way of disposing of polymer waste.

(1)

(Total for Question 4 = 10 marks)

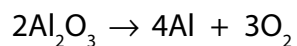


Metals

- 5 (a) (i) Complete the sentence by putting a cross (☒) in the box next to your answer. (1)

Aluminium is extracted from aluminium oxide by electrolysis.

The overall equation for this reaction is



In this reaction

- A aluminium oxide is oxidised
- B aluminium oxide is reduced
- C aluminium metal is reduced
- D oxygen gas is oxidised

- (ii) Part of the reactivity series is shown

aluminium	↑	most reactive
iron		
lead		
copper		least reactive

Lead is extracted from its oxide by heating the oxide with carbon rather than by using electrolysis.

Explain why.

(2)

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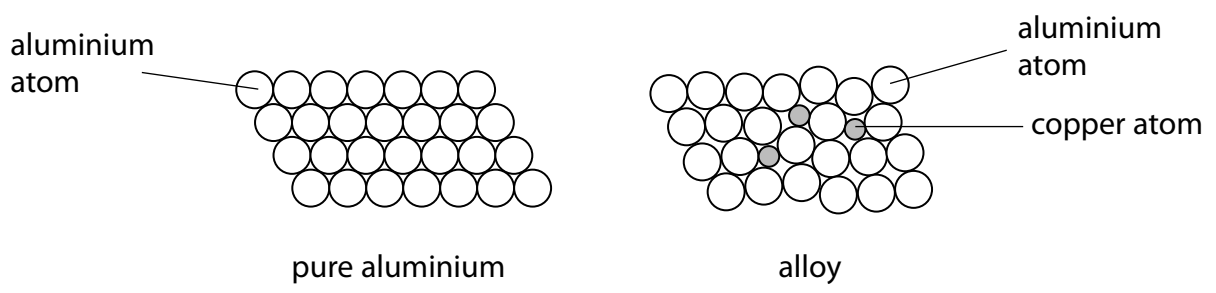
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(b) Pure aluminium is too weak for many uses.
Copper can be mixed with aluminium to produce an alloy.
The alloy is stronger than pure aluminium.

The diagrams show the structures of pure aluminium and the alloy.



Explain, in terms of these structures, how the presence of copper atoms in the alloy results in the alloy being stronger than pure aluminium.

(3)

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Handwriting practice area with 20 horizontal dotted lines.

(Total for Question 5 = 12 marks)



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Products from crude oil

6 Crude oil is separated into different fractions by fractional distillation.

- (a) Fraction X is obtained from near the top of the fractionating column.
 Fraction Y is obtained from near the bottom of the fractionating column.

Which row of the table shows the boiling point, ease of ignition and viscosity of fraction X compared with fraction Y?

Put a cross (☒) in the box next to your answer.

(1)

	boiling point	ease of ignition	viscosity
<input type="checkbox"/> A	lower than Y	more difficult than Y	higher than Y
<input type="checkbox"/> B	lower than Y	easier than Y	lower than Y
<input type="checkbox"/> C	higher than Y	more difficult than Y	lower than Y
<input type="checkbox"/> D	higher than Y	easier than Y	higher than Y

- (b) Pentane, C_5H_{12} , can be obtained from crude oil.
 When pentane burns completely in air, it forms carbon dioxide and water.

Write the balanced equation for this reaction.

(3)

- (c) Sulfur is present as an impurity in some fuels.

Explain how the product of combustion of this impurity in a fuel can damage the environment.

(2)

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(Total for Question 6 = 12 marks)

TOTAL FOR PAPER = 60 MARKS



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